

AD 744153

FTD-HT-23-1099-71

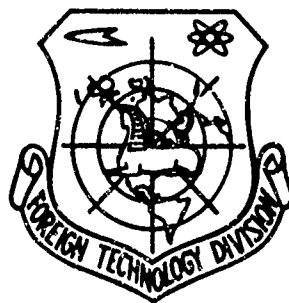
FOREIGN TECHNOLOGY DIVISION



TEMPERATURE DEPENDENCES OF THE VISCOSITY OF METALS

by

M. P. Bogdanovich
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UNCLASSIFIED

Security Classification

DOCUMENT CONTROL DATA - R & D

(Security classification of title, body of abstract and indexing classification must be entered when the overall report is classified)

1. ORIGINATING ACTIVITY (Corporate author) Foreign Technology Division Air Force Systems Command U. S. Air Force		30. REPORT SECURITY CLASSIFICATION UNCLASSIFIED	
		31. GROUP	
2. REPORT TITLE TEMPERATURE DEPENDENCES OF THE VISCOSITY OF METALS			
4. DESCRIPTIVE NOTES (Type of report and inclusion dates) Translation			
5. AUTHOR(S) (First name, middle initial, last name) Bogdanovich, M. P. and Baum, B. A.			
6. REPORT DATE 1968	7A. TOTAL NO. OF PAGES 4	7B. NO. OF REFS 10	
8A. CONTRACT OR GRANT NO.		8B. ORIGINATOR'S REPORT NUMBER(S) FTD-HT-23-1099-71	
8C. PROJECT NO.		8D. OTHER REPORT NO(S) (Any other numbers that are assigned this report)	
8E. DIA Task No. T69-04-9			
10. DISTRIBUTION STATEMENT Approved for public release; distribution unlimited.			
11. SUPPLEMENTARY NOTES		12. SPONSORING MILITARY ACTIVITY Foreign Technology Division Wright-Patterson AFB, Ohio	
13. ABSTRACT For pure metals three types of given dependences are observed: 1) the polytherm has its convex side facing the abscissa, which is probably explained by the weakening of bonds during the thermal expansion of a fluid; 2) the curve has its concave side to the abscissa, the activation energy increases with a rise in temperature, which can be caused by an increase in the coordination number, the integral bond energy increases, regardless of the possible weakening of a single bond, this can also take place in the case of intensive metallization of primarily covalent bonds; 3) the given linear dependence, the invariability of E with a rise in temperature is probably explained by the compensating effect of several factors.			

UNCLASSIFIED
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14. KEY WORDS	LINK A		LINK B		LINK C	
	ROLE	WT	ROLE	WT	ROLE	WT
Liquid Metal						
Temperature Dependence						
Gallium						
Copper						
Iron						
Cobalt						
Nickel						
Lithium						
Silicon						
Zinc						
Cadmium						

UNCLASSIFIED
Security Classification

EDITED TRANSLATION

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METALS

By: M. P. Bogdanovich, and B. A. Baum

English pages: 4

Source: Ural'skiy Politekhnicheskiy Institut.
Trudy, (Ural Politechnical Institute,
Transactions), No. 167, 1968, pp 96-98.

Translated by H. Peck

UR/0000-68-000-167

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PREPARED BY:

TRANSLATION DIVISION
FOREIGN TECHNOLOGY DIVISION
WP-APB, OHIO.

FTD-HT-23-1099-71

Date 13 Mar 19 72

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Block	Italic	Transliteration	Block	Italic	Transliteration
А а	<i>А а</i>	A, a	Р р	<i>Р р</i>	R, r
Б б	<i>Б б</i>	B, b	С с	<i>С с</i>	S, s
В в	<i>В в</i>	V, v	Т т	<i>Т т</i>	T, t
Г г	<i>Г г</i>	G, g	У у	<i>У у</i>	U, u
Д д	<i>Д д</i>	D, d	Ф ф	<i>Ф ф</i>	F, f
Е е	<i>Е е</i>	Ye, ye; E, e*	Х х	<i>Х х</i>	Kh, kh
Ж ж	<i>Ж ж</i>	Zh, zh	Ц ц	<i>Ц ц</i>	Ts, ts
З з	<i>З з</i>	Z, z	Ч ч	<i>Ч ч</i>	Ch, ch
И и	<i>И и</i>	I, i	Ш ш	<i>Ш ш</i>	Sh, sh
Й й	<i>Й й</i>	Y, y	Щ щ	<i>Щ щ</i>	Shch, shch
К к	<i>К к</i>	K, k	Ъ ъ	<i>Ъ ъ</i>	"
Л л	<i>Л л</i>	L, l	Ы ы	<i>Ы ы</i>	Y, y
М м	<i>М м</i>	M, m	Ь ь	<i>Ь ь</i>	'
Н н	<i>Н н</i>	N, n	Э э	<i>Э э</i>	E, e
О о	<i>О о</i>	O, o	Ю ю	<i>Ю ю</i>	Yu, yu
П п	<i>П п</i>	P, p	Я я	<i>Я я</i>	Ya, ya

* ye initially, after vowels, and after ъ, ь; e elsewhere.
 When written as ѣ in Russian, transliterate as yě or ě.
 The use of diacritical marks is preferred, but such marks
 may be omitted when expediency dictates.

FOLLOWING ARE THE CORRESPONDING RUSSIAN AND ENGLISH
DESIGNATIONS OF THE TRIGONOMETRIC FUNCTIONS

Russian	English
sin	sin
cos	cos
tg	tan
ctg	cot
sec	sec
cosec	csc
sh	sinh
ch	cosh
th	tanh
cth	coth
sch	sech
csch	csch
arc sin	sin ⁻¹
arc cos	cos ⁻¹
arc tg	tan ⁻¹
arc ctg	cot ⁻¹
arc sec	sec ⁻¹
arc cosec	csc ⁻¹
arc sh	sinh ⁻¹
arc ch	cosh ⁻¹
arc th	tanh ⁻¹
arc cth	coth ⁻¹
arc sch	sech ⁻¹
arc csch	csch ⁻¹
<hr/>	
rot	curl
lg	log

TEMPERATURE DEPENDENCES OF THE VISCOSITY OF LIQUID METALS

M. P. Bogdanovich and B. A. Baum

The activation energy of viscous flow E with a temperature invariability of the flow units is defined by the expression [1].

$$E = \frac{R}{\lg e} \cdot \frac{d(\lg \nu)}{d\left(\frac{1}{T}\right)},$$

where ν is the coefficient of kinematic viscosity.

For pure metals three types of dependences of $\lg \nu$ on T^{-1} are observed:

1. The polytherm has its convex side facing the abscissa, which is probably explained by the weakening of bonds during the thermal expansion of a fluid.

2. The curve has its concave side to the abscissa. The activation energy increases with a rise in temperature, which can be caused by an increase in the coordination number. The integral bond energy increases, regardless of the possible weakening of a single bond. This can also take place in the case of intensive metallization of primarily covalent bonds.

3. The linear dependence of $\lg \nu$ on T^{-1} . The invariability of

E with a rise in temperature is probably explained by the compensating effect of several factors.

A combination of the above dependences is also possible.

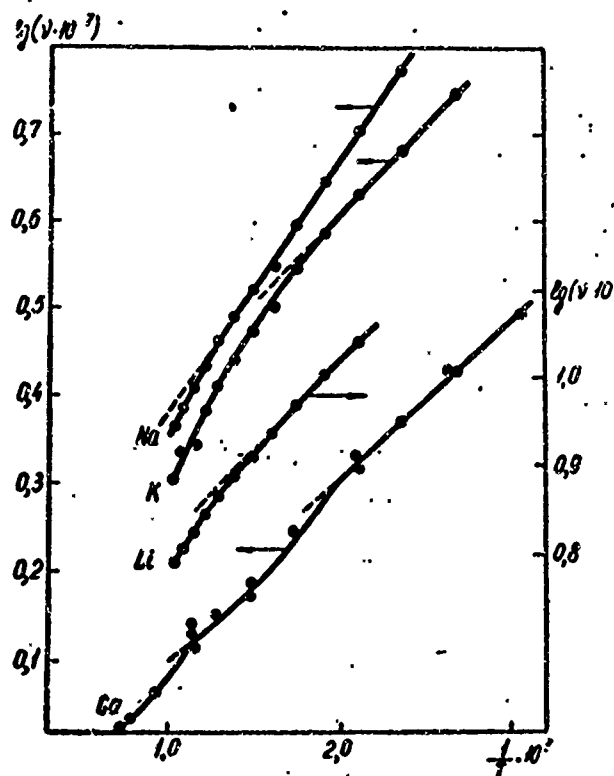
The first type is characteristic for Cu, Ag, and to a certain extent for Cr, Fe, Co, Ni (for example, [1, 2], and the second is characteristic for Li, Na, K, Si [1, 3], while the third is characteristic for Zn, Cd, Hg, Pb, Sn, Bi [3, 4].

For alkaline metals Li, Na, K (see the figure) the experimental shape of the temperature dependence of η is explained by an increase in the coordination number, which is confirmed by the data of X-ray structural analysis [3, 5]. The heating of liquid silicon is obviously accompanied not only by an increase in the coordination number but also by the decay of complexes with permanent bonds and the metallization of the bonds that occurs in connection with this [1]. In this case the dimensions of the flow units are changed, and the value of E calculated from equation (1) does not reflect the true energy value of the interparticle interaction in the melt.

The constancy of E in such metals as Pb and Zn cannot be explained only by an increase in the coordination number, because in the liquid state it is also quite high at $T_{\text{пл}}$. We can assume that in these metals the magnitude of the charge of the ions participating in the viscous flow increases, the dimensions of the ions decrease, and their mobility increases. This concept is confirmed to a certain extent by data on the temperature coefficient of electrical resistance $\beta = d\rho/dT$. Thus, for Zn, Cd, Hg, In, Pb, Sn and Bi, β in the liquid state ($\beta_{\text{л}}$) is less than in the solid state ($\beta_{\text{тв}}$) [6, 7]. At the same time, in Na, K, Cu, and Ag [8, 9] $\beta_{\text{л}} > \beta_{\text{тв}}$.

It is interesting to note that the temperature dependence of the viscosity of liquid gallium [10] is of a more complex (mixed) nature (see the figure). Up to 230°C it is linear. This, as in the preceding case, can be explained by the metallization of the bonds and by a certain increase in the coordination number. At temperatures of 230 and 560°C there are observed bends corresponding to the transfer to

the second type of dependence. Since the existence of groupings with primarily covalent bonds [5] is not excluded in gallium, we can assume that they are most intensely destroyed here.



Dependence of $\lg(\nu \cdot 10^7 \text{ m}^2/\text{s})$ on $1/T$ for certain liquid metals.

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